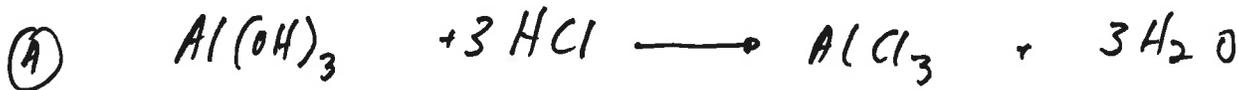


3.63 0.500g 0.702g = 0.855g 0.346g



(B) $0.500g Al(OH)_3 \times \frac{1 \text{ mole } Al(OH)_3}{78g Al(OH)_3} \times \frac{3 \text{ mole } HCl}{1 \text{ mole } Al(OH)_3} \times \frac{36.5g}{1 \text{ mole } HCl} = 0.702g HCl$

(C)

$\frac{1 \text{ mole } AlCl_3}{1 \text{ mole } Al(OH)_3} \times \frac{133.34g}{1 \text{ mole } AlCl_3} = 0.855g AlCl_3$

(E)

$\frac{3 \text{ mole } H_2O}{1 \text{ mole } Al(OH)_3} \times \frac{18g}{1 \text{ mole } H_2O} = 0.346g H_2O$

0.500g $Al(OH)_3$
0.702g HCl

Reactants = 1.202

0.855g
0.346g

Products = 1.201

$1.202g \approx 1.201g$

MASS IS CONSERVED

