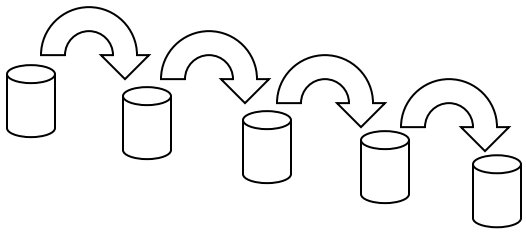
<u>Q101 Periodicity and Chemical Reactivity:</u>

This experiment determines what two variations in Group 2A?

What are all the members of Group 2A?

What would be your hypothesis?

Diagram how you would dilute a glass of Juice from full concentration to 1/16 concentration using the arrows and glasses below showing transfers between microplate wells and use ● to indicate the addition of water (10 drops = 0.5ml).



1	2	3	4	5	6	7	8	9

If the 1st well (A1) is 0.1M of solution, show the calculations that lead up to the last well (A9) being 4.8×10^{-5}M solution.

A1 = 0.1M solution
A2 = 0.05M solution
A3 =
A4 =
A5 = _____M solution
A6 = _____M solution
A7 = _____M solution
A8 = _____M solution
A9 = _____M solution

Mg^{+2}	1	2	3	4	5	6	7	8	9
Na_2S0_4									
Na_2CO_3									
$Na_2C_2O_4$									

Ca^{+2}	1	2	3	4	5	6	7	8	9
Na_2S0_4									
Na_2CO_3									
$Na_2C_2O_4$									

Sr^{+2}	1	2	3	4	5	6	7	8	9
Na_2S0_4									
Na_2CO_3									
$Na_2C_2O_4$									

Ba^{+2}	1	2	3	4	5	6	7	8	9
Na_2S0_4									
Na_2CO_3									
$Na_2C_2O_4$									

Figure 1: (provide a meaning full title for this figure: explain what is it? why did you do it? what were the results?)

↑P = high level precipitant
- P = precipitant
↓P = low level precipitant
☒ = no reaction

After reading exercise ***Dr. DUICK***

Complete the graphic organizer that depicts the order of the experiment:

Word bank:

Serial dilutions	Measure out initial 0.1M solutions
Record sodium oxalate results	Record sodium carbonate results
Record sodium sulfate results	Add Na_2S0_4 Add $Na_2C_2O_4$
Add Na_2CO_3	Hypothesis
Serial dilutions	Serial dilutions
Label Microplates	Conclusions